

# Edexcel Chemistry GCSE

## CP 2 - Investigating pH (neutralisation)

### Flashcards

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Volumetric Pipette

Greater accuracy than a measuring cylinder



# How can you measure the pH of a solution?



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Use a pH probe or add universal indicator and compare to a colour chart

- Red: strong acid (pH 1)
- Yellow: weak acid
- Green: neutral (pH 7)
- Light blue: weak alkali
- Blue/purple: strong alkali (pH 14)



# What is a neutralisation reaction?



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Hydrogen ions reacting with hydroxide ions to form water



When measuring the pH change of the combination of two reactants why is it important to stir the mixture after the reactants are combined?





When measuring the pH change of the combination of two reactants why is it important to stir the mixture after the reactants are combined?

To ensure a complete reaction and that the pH is consistent throughout the mixture



How could an experiment measuring the pH using universal indicator paper be improved?



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Use a pH probe to get an accurate reading



$\text{Ca(OH)}_2$  can be added to HCl for neutralisation. How do you know that the HCl has been fully neutralised?



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The pH should be pH 7 - neutral



Ammonia solution is an alkali. What could be used to show it is alkaline?



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Either:

- Universal indicator paper turns blue
- High pH probe value around 11

